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thermal reservoir at constant temperature  $T_h$ , converts an amount  $W$  of this heat into work, and expels the remainder  $Q_c$  into another thermal reservoir at a constant temperature  $T_c$ . The efficiency  $e$  of the heat engine is the fraction of  $Q_h$  that is converted into work, so  $e = \frac{W}{Q_h} = \frac{Q_h - Q_c}{Q_h} = 1 - \frac{Q_c}{Q_h}$  (1) Since the absorbed heat  $Q$

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This agrees with Schroeder's Figure 1.14. To put in some numbers, we can use the data from the earlier problem. For water, the values given by Schroeder are (at  $25^\circ\text{C} = 298\text{ K}$ ):  $C_V = 2.57 \times 10^4 \text{ J K}^{-1} \text{ (20)}$   $T = 4.52 \times 10^{10} \text{ Pa}^{-1} \text{ (21)}$  For one mole of water, the volume is  $18.068 \times 10^{-6} \text{ m}^3$ , so  $C_P - C_V = (298) 18.068 \times 10^{-6} \times 2.57 \times 10^4 = 1.38 \times 10^4 \text{ J K}^{-1} \text{ (22)}$   $= 0.787 \text{ J K}^{-1} \text{ (23)}$   $C$

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